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Gas Laws Worksheet

Boyle's Law Problems: $P_1V_1 = P_2V_2$

1. If 22.8 L of nitrogen at 740 mm Hg are compressed to 723 mm Hg at constant temperature, what is the new volume?

$$\frac{(740 \text{ mm Hg})(22.8 \text{ L})}{(723 \text{ mm Hg})} = V_2$$
$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$\frac{(205 \text{ kPa})(4.0 \text{ L})}{(12.0 \text{ L})} = P_2$$
$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.8 liters?

$$\frac{(1.00 \text{ atm})(196.0 \text{ L})}{(26.8 \text{ L})} = P_2$$
$$P_2 = 7.30 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$\frac{(12.7 \text{ kPa})(40.0 \text{ L})}{(8.4 \text{ kPa})} = V_2$$
$$V_2 = 60.5 \text{ L}$$

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