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Mole Conversions Worksheet Answers

Mole Conversions Worksheet

There are three mole equalities. They are:
1 mol = 6.02×10^{23} particles
1 mol = g/formula-mass (periodic table)
1 mol = 22.4 L for a gas at STP

Each equality can be written as a set of two conversion factors. They are:

$$\left(\frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ particles}} \right) \quad \left(\frac{6.02 \times 10^{23} \text{ particles}}{1 \text{ mole}} \right)$$
$$\left(\frac{1 \text{ mole}}{g - \text{formula-mass}} \right) \quad \left(\frac{g - \text{formula-mass}}{1 \text{ mole}} \right)$$
$$\left(\frac{1 \text{ mole}}{22.4 \text{ L}} \right) \quad \left(\frac{22.4 \text{ L}}{1 \text{ mole}} \right)$$

Mole-Particle Conversions

- How many moles of magnesium is 3.01×10^{22} atoms of magnesium?
 $3.01 \times 10^{22} \text{ atoms} \left(\frac{1 \text{ mole}}{6.02 \times 10^{23} \text{ atoms}} \right) = 5 \times 10^{-2} \text{ moles}$
- How many molecules are there in 4.00 moles of glucose, $C_6H_{12}O_6$?
 $4.00 \text{ moles} \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mole}} \right) = 2.41 \times 10^{24} \text{ molecules}$
- How many moles are 1.20×10^{22} atoms of phosphorus?
- How many atoms are in 0.750 moles of zinc?
- How many molecules are in 0.400 moles of N_2O_5 ?