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Covalent Bonding Review Sheet - Dec 17, 2012

- Name _____ Date period _____
- Given the formula, a compound can be named using conventional systems that include Greek prefixes and Roman numeral where appropriate.
 - Differences in electronegativity values can be used to predict where a bond fits on the continuum between ionic and covalent bonds.
 - Some most compounds contain multiple bonds, a substance may contain more than one type of bond.

Section 1: Naming and Formula Writing

Write the formula for the following covalent compounds

- Nitrogen Monoxide NO
- Dinitrogen Monoxide N_2O
- Nitrogen Dioxide NO_2
- Dinitrogen Trioxide N_2O_3
- Dinitrogen Tetraoxide N_2O_4
- Dinitrogen Pentoxide N_2O_5
- carbon dioxide CO_2
- phosphorus trioxide P_2O_3
- sulfur dichloride SCl_2
- nitrogen trifluoride NF_3

Write names from the following formula:

- NF_3 _____ dinitrogen tetrafluoride _____
- SCl_4 _____ sulfur tetrachloride _____
- CF_4 _____ carbon tetrafluoride _____
- SO_2 _____ sulfur dioxide _____
- P_2O_5 _____ diphosphorus pentoxide _____
- SF_6 _____ sulfur hexafluoride _____
- H_2O _____ hydrogen monoxide _____
- Br_2 _____ bromine _____ or bromine _____
- HBr _____ hydrogen monobromide _____
- N_2O_3 _____ dinitrogen trioxide _____

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